

表 1.1: H₂-O₂ 反応機構 (Chemkin 形式)

reactions			
! - major three chain branching/propagation reactions			
H+O2=OH+O	9.756e+013	0.000	14844.6
O+H2=OH+H	5.120e+004	2.670	6278.7
H2O+H=H2+OH	4.520e+008	1.600	18422.6
! - major near-termination forming HO2			
H+O2+M=HO2+M	2.100e+018	-0.800	0.0
N2/0.67/ O2/0.4/ CO2/1.5/ H2O/0/ AR/0.29/			
O2+H+H2O=HO2+H2O	6.890e+015	0.000	-2086.5
! - OH/HO2 radical chain degradation/termination			
OH+OH=O+H2O	1.510e+009	1.140	100.4
OH+HO2=H2O+O2	2.890e+013	0.000	-497.1
H+HO2=H2+O2	4.280e+013	0.000	1410.1
HO2+HO2=H2O2+O2	4.220e+014	0.000	11983.7
duplicate			
HO2+HO2=H2O2+O2	1.320e+011	0.000	-1630.0
duplicate			
! - HO2 radical chain reactivation/propagation			
H+HO2=OH+OH	1.690e+014	0.000	874.8
H+HO2=H2O+O	3.010e+013	0.000	1720.8
O+HO2=O2+OH	3.190e+013	0.000	0.0
! - H2O2 reactions			
OH+OH(+M)=H2O2(+M)	7.230e+013	-0.370	0.0
low / 5.530e+019 -0.760 0.0 /			
troe / 1 1 1 1040 /			
N2/0.4/ O2/0.4/ CO2/1.5/ H2O/6.5/ AR/0.35/			
H2O2+H=HO2+H2	1.690e+012	0.000	3754.8
H2O2+H=OH+H2O	1.020e+013	0.000	3577.9
H2O2+O=OH+HO2	6.620e+011	0.000	3974.7
H2O2+OH=H2O+HO2	7.830e+012	0.000	1331.3
! - recombination chain termination/degradation			
H+H+M=H2+M	1.870e+018	-1.000	0.0
N2/0.4/ O2/0.4/ CO2/1.5/ H2O/6.5/ AR/0.35/			
H+H+H2=H2+H2	9.790e+016	-0.600	0.0
H+O+M=OH+M	1.180e+019	-1.000	0.0
N2/0.4/ O2/0.4/ CO2/1.5/ H2O/6.5/ AR/0.35/			
H+OH+M=H2O+M	5.530e+022	-2.000	0.0
N2/0.4/ O2/0.4/ CO2/1.5/ H2O/2.55/ AR/0.15/			
O+O+M=O2+M	5.400e+013	0.000	-1787.8
N2/0.4/ O2/0.4/ CO2/1.5/ H2O/6.5/ AR/0.35/			
end			

表 1.2: Cl₂-H₂ 連鎖反応系の反応速度定数

反応	A [cm ³ mol ⁻¹ s ⁻¹]	E _a / R [K]	k (298 K) [cm ³ mol ⁻¹ s ⁻¹]
(1) Cl + H ₂ → H + HCl	2.2 × 10 ¹³	2300	9.8 × 10 ⁹
(2) H + Cl ₂ → Cl + HCl	4.8 × 10 ¹³	416	1.2 × 10 ¹³

表 1.3 H₂-O₂ 連鎖反応系の反応速度定数

反応	A [cm ³ mol ⁻¹ s ⁻¹]	b	E _a / R [K]	k (1000 K) [cm ³ mol ⁻¹ s ⁻¹]
(1) H + O ₂ → OH + O	9.8 × 10 ¹³	0	7470	5.6 × 10 ¹⁰
(2) O + H ₂ → OH + H	5.1 × 10 ⁴	2.67	3160	2.2 × 10 ¹¹
(3) OH + H ₂ → H ₂ O + H	1.0 × 10 ⁸	1.6	1660	1.2 × 10 ¹²
* (4) H + O ₂ + M → HO ₂ + M	1.7 × 10 ¹⁸	-0.8	0	6.8 × 10 ¹⁵

* (4) の A と k (1000K) の単位は [cm⁶ mol⁻² s⁻¹].

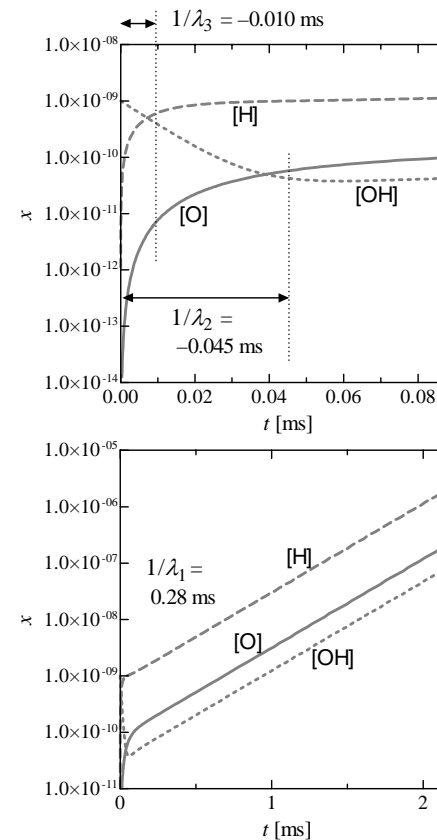


図 1.1: H₂-O₂ 連鎖反応
[H₂:O₂ = 2:1, p = 0.01 atm,
T = 1000 K, x(OH)₀ = 1 × 10⁻⁹]

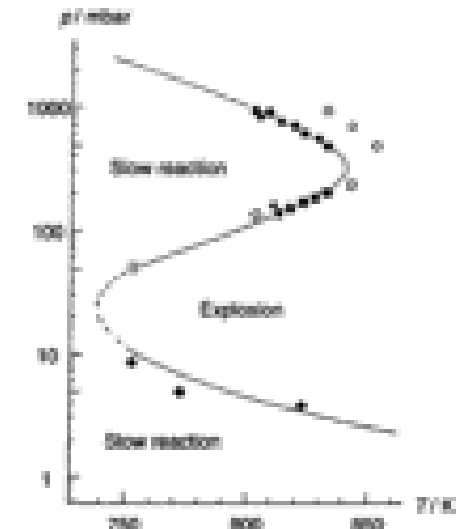


図 1.2: H₂-O₂ 爆発限界
800 K 付近を見たときに限界曲線を
下から (低圧から) 順に、第一限界、
第二限界、第三限界と呼ばれる。